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Research Article

Study of Hydrolysis Process from Pineapple Leaf Fibers using Sulfuric Acid, Nitric Acid, and Bentonite Catalysts

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Abstract

The hydrolysis process of pineapple leaf fibers has been carried out using sulfuric acid, nitric acid, bentonite catalyst, and activated bentonite catalyst. The sugar content of the hydrolysis product was identified using the phenol-sulfuric acid method by UV-Visible spectrophotometry. The disposal of pineapple leaf is a big problem even though it has high cellulose content (70–80%) and very promising to produce sugar by hydrolysis process. The purpose of this experiment was to determine the effectiveness of homogeneous and heterogeneous catalysts related to sugar levels in pineapple leaf fiber. The variables in this study were the sampling time during the hydrolysis process at a temperature of 100 °C and the addition of homogeneous and heterogeneous catalysts. The homogeneous catalysts were sulfuric acid and the nitric acid meanwhile heterogeneous catalyst was thermally activated bentonite and acid-activated bentonite. The results obtained highest sugar content reached at 150 minutes using chemical activated bentonite catalysts at 6.459 g/L and the addition of catalysts affected sugar yields, speed up the reaction, bentonite as a good catalyst, and gave good prospect for ethanol production in further process.

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Keywords: pineapple leaf fiber; hydrolysis; bentonite catalysts; kinetics

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1. Introduction

Pineapple (*Ananas comosus* L. Merr) is one type of fruit whose marketing is fairly evenly distributed in regions of Indonesia. Pineapple production in Indonesia amounted to 1,882,806 tons and the focus of plant cultivation is from the fruit while the other parts become the residual waste, such as: the leaf, pulp, skin, and peel. Pineapple skin waste produced from one pineapple fruit ranges from 21.73–24.48%, so that the pineapple skin waste produced can reach 40-50 kg/day. Pineapple leaf is used for animal feed

although it is not known considered due to high fiber content, high soluble carbohydrate, and low protein content [1]. Pineapple leaves are the most widely produced waste from pineapple farming, which is around 90% of each harvest, and each time a pineapple harvest produces waste consisting of 1% stems, 9% stem shoots, and 90% leaves. This becomes a consideration of utilizing pineapple leaf as a lignocellulose source that other parts of residues. Pineapple leaf contain 69.5–71.5% cellulose and 4.4–4.7% lignin [2], while there are 69.5–71.5% cellulose in pineapple leaf fiber [3].

Pineapple leaf fiber waste has not been optimally used although it was a potential biomass source due to high cellulose content for generat-

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ing reducing sugar through a hydrolysis process to be fermented into ethanol or other value-added products [4]. Cellulose can be hydrolyzed into sugar monomers both chemically with acid solution and enzymatic compounds with cellulase enzyme [5]. Both of these technologies that are often used still have technical constraints namely the price of enzymes that are expensive and vulnerable to changes in operating conditions. While acid hydrolysis is corrosive and causes waste, so it is needed to develop new technologies to improve technical problems, one of which is the heterogeneous catalyst.

The process of hydrolysis made from lignocellulose that has been carried out includes hydrolysis of water hyacinth and rice husk using sulfuric acid to produce reducing sugar 23.3 mg/g [6]. Pineapple leaves with hydrochloric acid (HCl) catalyst of 0.3N produce the highest glucose levels, 8.958–9.594% [7]. Hydrolysis of weeds (*Imperata cylindrica*) with sulfonated activated carbon catalyst resulted in a glucose level of 87.2% [8], while another report used similar substrate hydrolyze using 2% sulfuric acid at 140 °C generated 5.0675% alcohol [9]. The main advantages of hydrolysis with dilute acid were no need for acid recovery and no loss of acid in the process [10]. The acid used is generally H_2SO_4 or HCl in the 2–5% concentration range [11].

In this study, pineapple leaf fiber was used in the hydrolysis process using the refluxing method with the assistance of a homogeneous and heterogeneous catalyst for the conversion of lignocellulose into total sugar. The results of

this study are expected to be used as a basis for developing pineapple leaf fiber into total sugar and figure out the effectiveness of sulfuric acid, nitric acid, and bentonite as a catalyst in hydrolysis processes based on the total sugar concentration and rate of reaction.

2. Materials and Methods

2.1 Materials

All reagents used were analytical grades from Merck such as concentrated sulfuric acid, concentrated nitric acid, phenol, and glucose standard. Pineapple leaves as raw material, and sodium bentonite from PT Tunas Inti Makmur Semarang. The equipment used was a set of reflux equipment, oven (Memmert UN55), furnace (Themolyne), analytical balance (Ohaus PA214), glassware, and UV-Visible spectrophotometer Thermo scientific GENESYS 20. The scheme of the hydrolysis process is described in Figure 1.

2.2 Sample Preparation

Pineapple leaf fiber weighed 20 grams added 200 mL distilled water and added to the chemical activated natural bentonite weighing 2.5 grams. The sample was put into a three-neck flask and refluxed for 3 hours at 100 °C. Variations of time sampling during the hydrolysis process are 30, 60, 90, 120, and 150 minutes. The hydrolysis solution was taken and analyzed at glucose level using the UV-Visible spectrophotometer at wavelengths of 480 nm and 490 nm. The same procedure was repeated for the hydrolysis process using 18% of H_2SO_4 and 12% of HNO_3 , no catalyst as a control system, and bentonite for both physical and chemical treatment. The activation of solid catalysts respectively was bentonite put into a furnace for 4 hours at 400 °C and refluxing bentonite in 18% of H_2SO_4 .

2.3 Analysis of Total Sugar

The preparation of a standard glucose solution is made by diluting the 100 mg/L glucose mother solution to a concentration of 0.005, 0.015, 0.020, 0.030, 0.040 and 0.05 mg/mL. Furthermore, each 1 mL standard solution was added 1 mL 5% phenol, shaken, and added 5 mL of concentrated sulfuric acid solution quickly by pouring perpendicularly to the surface of the solution. The solution is allowed to stand for 10 minutes, beaten then placed in a bath filled with warm water for 15 minutes. Absorbance measurements were performed us-

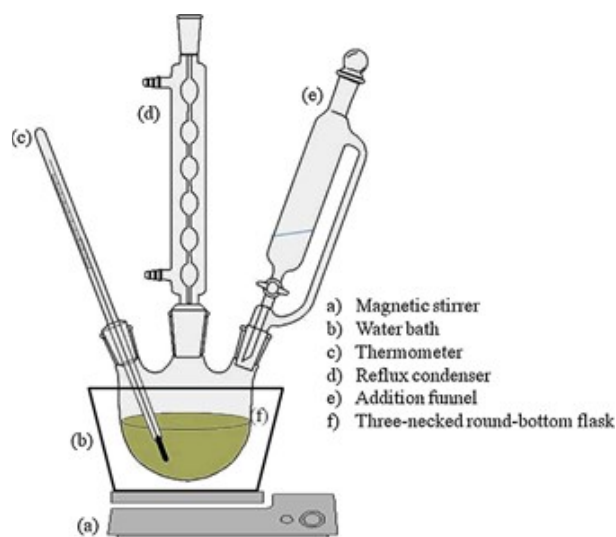


Figure 1. Scheme of hydrolysis process using reflux.

ing a UV-Visible spectrophotometer at 480 nm and 490 nm.

The 1 mL hydrolysate solution and 1 mL phenol 5% were shaken then added 5 mL of concentrated sulfuric acid quickly by pouring perpendicularly to the surface of the solution. The solution is allowed to stand for 10 minutes and shake again before placed in a water bath for 15 minutes. Absorbance measurements were performed using a UV-Visible spectrophotometer at a similar wavelength as glucose standard solution [12].

2.4 Kinetic Study

The kinetic study was measured based on experimental results of the various time of the hydrolysis reaction (30, 60, 90, 120, and 150 minutes). Based on these data, the studies of kinetic were calculated by zeroth-order reaction, 1st order reaction, and 2nd order reaction models.

3. Results and Discussion

3.1 Analysis of Total Sugar

Almost all types of natural fibers, specifically derived from plants (vegetable fibers) like pineapple leaves have the biggest chemical fiber is cellulose, even though it is an element or other substances also found in these fibers, for example, fats and wax, hemicellulose, lignin, and pectin. Fiber obtained from young pineapple leaves has relatively low strength and shorter fiber compared to fiber from old leaves. Pineapple fiber consists of cellulose and non-cellulose obtained through mechanical removal of the outer layer of leaves. Cellulose is one of the chemical compositions in pineapple leaves that has the highest content. Cellulose ($C_6H_{10}O_5$)_n is a long-chain polymer of carbohydrate polysaccharides, from β -glucose. Cellulose is one of the main components of lignocellulose which consists of D-glucose monomer units that are bound to 1,4-glycosidic bonds. Cellulose

tends to form microfibrils through inter and intramolecular bonds thus providing a soluble structure. The chemical composition of pineapple fiber is listed in Table 1.

About 69–71% of cellulose found in pineapple fiber was the main source of sugar production which through enzymatic follow-up treatment will produce ethanol. The abundant cellulose fiber in pineapple leaves will be hydrolyzed in a dilute acid solution or with the help of a solid acid catalyst while pentosan (which is part of hemicellulose) and lignin will inhibit the hydrolysis reaction because resistant to chemicals. Total sugar that was produced from the hydrolysis process of cellulose and hemicellulose can be measured using various methods that are frequently performed by several colorimetric procedures. Most methods involve colorimetric detection based on the oxidation of the carbonyl groups followed by reaction with a UV-Visible absorbing molecule. The well-known method was phenol-sulfuric acid for determining sub-micro amounts of sugars and related substances due to good sensitivity, low analysis cost, easy handling, and accessible equipment although it requires a high concentration of hazardous sulfuric acid and a wide range of analytical wavelengths for measurement. This method used a 5% phenol solution and concentrated sulfuric acid which is added after the hydrolysis results of pineapple leaf fibers and after 15 minutes of cooling into room temperature to help the process of color formation (complexing reaction) becoming stable as described in Figure 2.

The sulfuric acid causes the hydrolysis of the glycosidic links and the dehydration of the released monosaccharides to form furfural derivatives which react with the phenol and give the orange color to the solution. The total sugar content from pineapple leaf fiber was quite high so it resulted in purple color as the final solution. The color produced by furfural methyl hydroxy varies from green, orange, brown, and

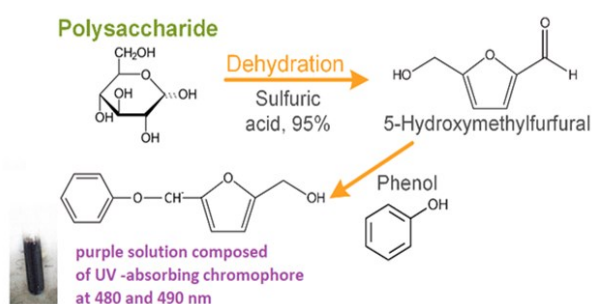


Figure 2. Schematic overview of total sugar analysis (modified from [11]).

Table 1. Chemical composition of pineapple fiber [3].

Composition	Amount (%)
α -cellulose	69.5–71.5
Pentosan	17–17.8
Lignin	4.4–4.7
Pectin	1–1.2
Fat and wax	3–3.3
Ash	0.71–0.87
Other substances (protein, organic acid, etc.)	4.5–5.3

purple depending on the concentration of glucose possessed by the sample. The measurement of hydrolysate products of pineapple leaf fiber at two different wavelengths due to each sugar has its maximum absorption related to the structure it is own. Each hydrolysate from the hydrolysis process with the addition of homogeneous and heterogeneous catalysts was calculated by extrapolating the absorbance value into the regression equation from the standard series of glucose as presented in Figure 3.

For the hydrolysis of pineapple leaf fibers using homogeneous catalysts, the total sugar content is different from heterogeneous catalysts where sulfuric acid catalysts can compete with heterogeneous catalysts compared to nitric acid catalysts. For the duration of the hydrolysis time of 150 minutes, the sulfuric acid catalyst was able to compete with a chemical activated catalyst with a total sugar content of 6.387 g/L even though chemical activated catalyst appeared dominant until the end of the hydrolysis time duration with a total sugar acquisition of 6.459 g/L. Optimal levels of total sugar of pineapple leaf fibers with a sulfuric acid catalyst and nitric acid catalyst at 150 minutes and did not increase significantly from 120 minutes to 150 minutes caused by degradation to side products such as furfural, 5-hydroxymethylfurfural (HMF), levulinic acid, acetic acid, formic acid, uronic acid, and so on.

These results are consistent with the previous study that hydrolysis time besides acid concentration and hydrolysis temperature plays an important role in the acquisition of total sugar content (in this case glucose) the hydrolysis of pineapple leaves [14]. Sulfuric acid catalysts produce more total sugar levels because it has higher acidity than nitric acid which is related to the value of acid dissociation constant where $pK_a \text{ HNO}_3 > pK_a \text{ H}_2\text{SO}_4$ made sulfuric acid could contribute more protons in aqueous solution. At the same hydrolysis time and concentration, sulfuric acid can break the hydrogen bonds until it reaches the amorphous and crystalline cellulose networks into oligosaccharide and monosaccharide molecules [15]. This is in line with the results of the hydrolysis of sago pulp waste using a sulfuric acid catalyst capable of producing a greater percentage of cellulose (59.51%) with a small difference in the percentage of yield (42.88%) compared to a stronger chloride acid catalyst (42.91%) [16]. Total sugar levels obtained from the use of sulfuric acid and nitric acid catalysts have decreased with the length of hydrolysis due to the process of sugar degradation and the formation of side products, such as: furfural, 5-hydroxymethylfurfural (HMF), levulinic acid, acetic acid, formic acid, uronic acid, and so on [17]. These results are in line with the previous report that using pineapple leaves for the hy-

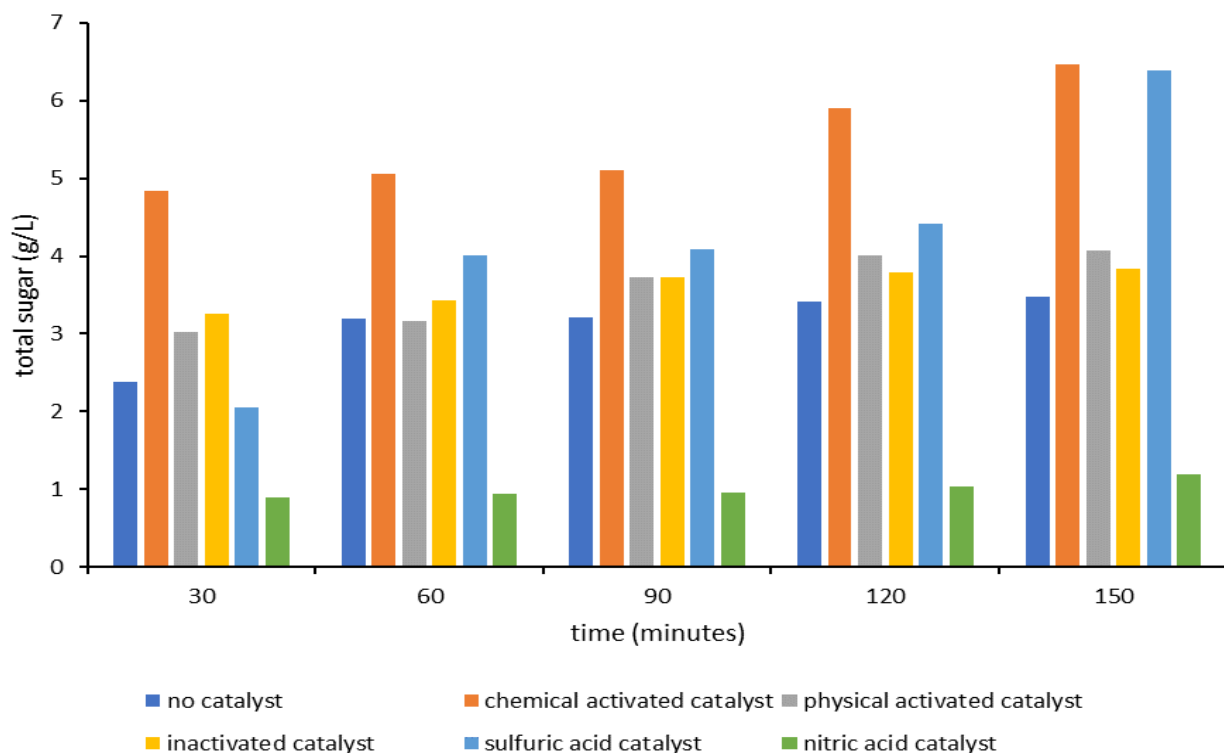


Figure 3. The total sugar content of hydrolysate.

hydrolysis process with a sulfuric acid solution at various concentrations (1–5%) and hydrolysis time (30–120 minutes) [18] and also the usage of catalyst HCl solution (0.1–0.3N) at temperatures of 30 °C and 80 °C up to 360 minutes for hydrolysis of pineapple skin [7].

Meanwhile, for the hydrolysis process using heterogeneous catalysts in the form of natural bentonite, physically activated bentonite, and chemical activated bentonite give good results and can compete with homogeneous catalysts. The hydrolysis process without the presence of a maximum catalyst produced a total sugar level of 3.48 g/L for 150 minutes which was able to obtain natural clay catalysts (3.43 g/L) with a duration of hydrolysis time of 60 minutes.

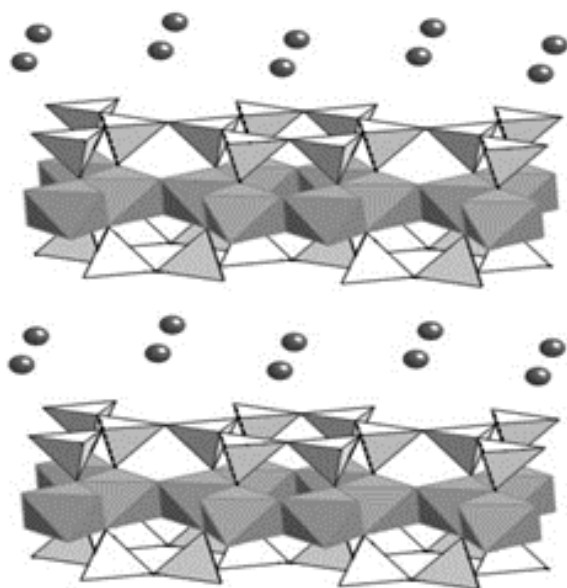


Figure 4. Structure of montmorillonite [19].

These results indicate that the catalyst solids help the hydrolysis process of pineapple leaf cellulose fibers become faster due to the acidity characteristics possessed by the catalyst. The heterogeneous catalyst used is the bentonite in a 2:1 configuration that has each aluminum-containing layer covalently bound above and below by a silicon-containing layer; this prevents active sites from interacting with the cellulose. The octahedra are AlO_6 units and the tetrahedra are SiO_4 units. The sheets in bentonite are charged due to Al:Si imbalance. The sheets are held together by ionic forces between the sheets and a cation (dark sphere) as described in Figure 4.

The process of cellulose hydrolysis in pineapple leaf fibers using a solution of sulfuric acid and nitric acid through the cleavage mechanism of β -1-4-glycosidic bonds takes place in three stages: the protons in the acid solution interact rapidly with oxygen (in the glycosidic bonds) to form acids the conjugate which then occurs by breaking the CO bonds and the conjugate acid into a cyclic carbonium ion as presented in Figure 5. The addition of fast water solvents in large quantities will release sugars and protons, where the process of breaking the glycosidic bonds is faster at the ends than in the middle of the polysaccharide chain [20]. Sulfuric and nitric acid catalysts used in this study were 18% and 11% respectively so that the rate of hydrolysis of pineapple leaf cellulose was slow because the catalyst concentration factor was very influential where concentrations above 65% were able to produce a hydrolysis rate twice as large so that it could produce a more optimal total sugar content. The water solvent hydrolysis system without pH control makes the water act as a weak acid which

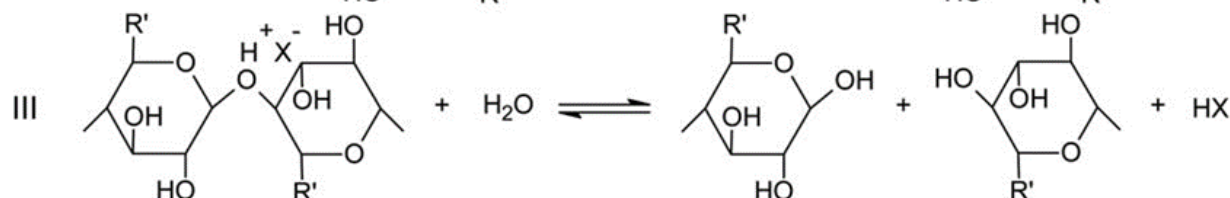
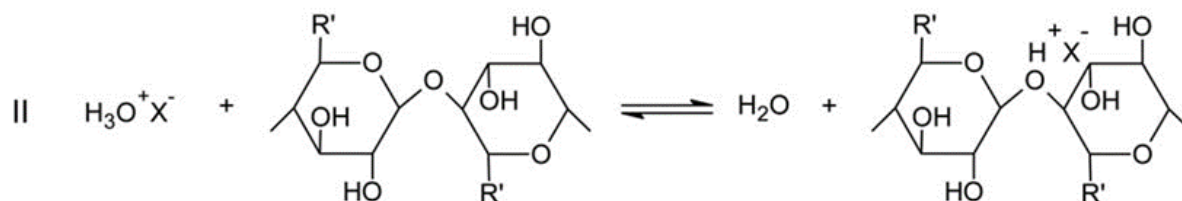


Figure 5. The mechanism of homogeneous acid-catalyzed cellulose hydrolysis [21].

quickly releases protons for the hydrolysis of polysaccharides into monosaccharides.

The acidity of bentonite clay is obtained from the active sites of Brönsted and Lewis which play a role in the hydrolysis of pineapple leaf cellulose fibers. To increase the catalytic ability of the bentonite catalyst the physical activation step was carried out through a calcination process at 600 °C for 4 hours and chemical activation using sulfuric acid. High-temperature heating increases the number of acid sites due to the formation of unsaturated Al^{3+} ions for Lewis acid and H_3O^+ ions which combine with AlO_4 in the tetrahedral layer where the highest amount is obtained when heating temperatures to 120 °C as reported by [22]. Meanwhile, chemical activation with sulfuric acid causes balancing cations, such as Na^+ , disappears and the structural changes occur because the number of dissolved Al ions makes the acidity level is low to moderate [23]. Table 2 showed the acidity value of bentonite catalyst was measured using the old fashion method which is ammonia-gravimetric. It is seen that activation of bentonite catalyst can increase the acidity and further enhance the catalytic process of cellulose hydrolysis.

Figure 6 presented the shifted band of infrared spectrum from raw and treated bentonite. Acid treatment on chemical activated bentonite has given a significant effect on remove water from the structure indicated by a less intense band at 3630 cm^{-1} and 1638 cm^{-1} also the band at 3435 cm^{-1} disappeared. These phenomena also emerge on treated bentonite by heating or physical process. There is also a decrease in the intensity of the band at 524 cm^{-1} (Si–O deformation) of treated samples compared to raw bentonite showed the affected of a tetrahedral

structure by acid treatment [23]. The bands around 3400 cm^{-1} for all samples were still seen indicated adsorbed water was desorbed by NH_3 adsorption, where interlayer water reacted with NH_3 forming NH_4^+ ions or also reacted with Lewis and Brönsted acid sites. The characteristic bands of hydrogen-bond ammonia were not observed due to overlap with the bands of NH_4^+ and NH_3 [22].

In general, the acid activation and heating effect increase the specific surface area of the bentonite. Hydrolysis of bentonite with acids disrupt the layered structure of the bentonite due to the aluminum oxide layer is attacked so that it dissolves some aluminum and other elements in the octahedral layer. The release of Al octahedral cations from the bentonite structure opens the structure with a distribution of one sheet of pore size, so that the relative acidity of the bentonite increases.

Brönsted acid sites originate from protons bound to siloxane (Si–O–Si) groups, while Lewis acid comes from Al metals from the octahedral groups that make up natural bentonite. The bentonite catalyst in this study still needs further comprehensive characterization to support acidity information as well as structural changes. Otherwise, the bentonite catalyst supported with Lewis acid improves the cellulose degradation due to Al as transition metal has numerous empty orbitals which can be occupied by lone pair electrons from oxygen in hydroxyl of cellulose generating the change in reaction mechanism as well as reducing the reaction of energy barrier as shown in Figure 7. A similar mechanism also appeared in a previous study using CrCl_3 catalyst for degradation of cellulose into glucose and HMF in variation temperature among 140 °C to 170 °C [24].

Every chemical reaction generates a specific chemical rate which can be a slow or fast reaction and catalysts are one of the factors affecting the chemical rate. The rate measurement based on a relationship between physical properties (λ) and concentration is described in Equations (2) and (4). In cellulose hydrolysis of pineapple leaf fiber, the addition of homogeneous and heterogeneous catalysts should be de-

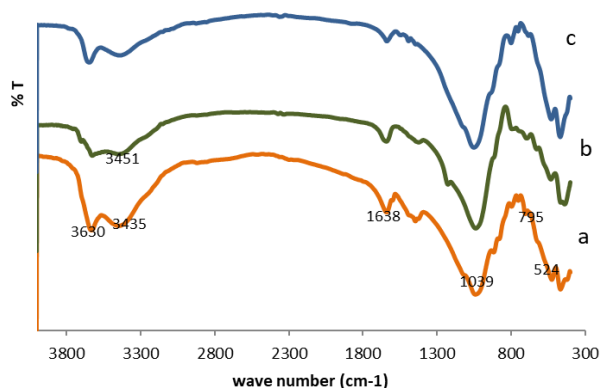


Figure 6. The infrared spectrum of (a) raw bentonite (b) physical activated bentonite and (c) chemical activated bentonite.

Table 2. Acidity value of the catalyst.

Catalyst	Acidity value (mmol/g)
Bentonite	5.122702
Physically activated bentonite	6.879196
Chemical activated bentonite	7.340855

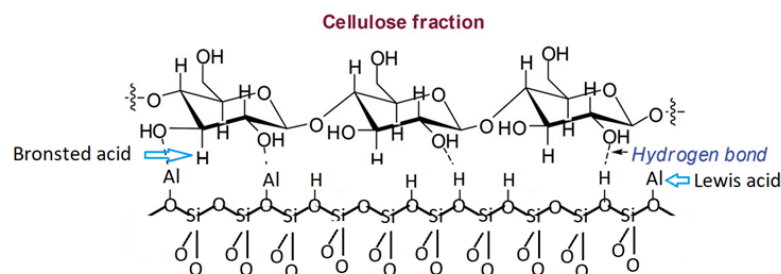


Figure 7. The mechanism of heterogeneous acid-catalyzed cellulose hydrolysis.

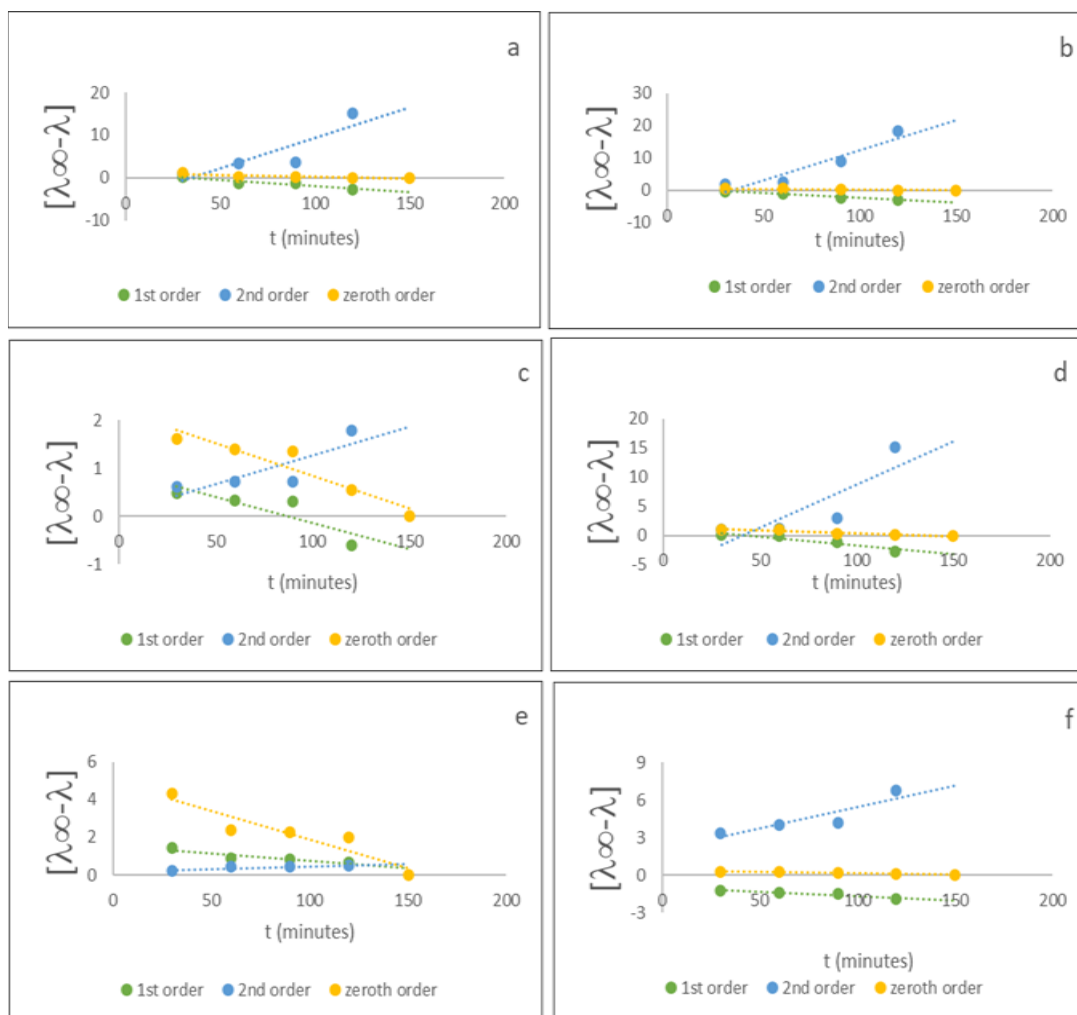


Figure 8. Graphical plot of rate constant at 100 °C for hydrolysis process with (a) no catalyst (b) inactivated catalyst (c) physical activated catalyst (d) chemical activated catalyst (e) sulfuric acid catalyst (f) nitric acid catalyst.

Table 3. Linearity value from the kinetic model.

Type of catalyst	R ²		
	Zeroth order reaction	1 st order-reaction	2 nd order-reaction
No catalyst	0.7584	0.9158	0.7624
Inactivated bentonite	0.9084	0.9579	0.8965
Physically activated bentonite	0.8915	0.7350	0.6818
Chemical activated bentonite	0.9307	0.8826	0.7100
Sulfuric acid catalyst	0.8672	0.8506	0.7948
Nitric acid catalyst	0.8804	0.8657	0.8122

terminated using the zeroth-order, 1st order, and 2nd order models as described respectively in Equations (1), (2), and (3) with λ_∞ are physical properties at an infinite time; λ_t are physical properties at $t=t$; λ_0 is physical properties at $t=0$; t is a time of reaction, A_0 is initial concentration; A_t is concentration at $t=t$; and k is the chemical reaction rate constant.

$$\lambda_t = -kt + \lambda_0 \quad (1)$$

$$\ln \left(\frac{\lambda_\infty - \lambda}{\lambda_\infty - \lambda_0} \right) = -kt \quad (2)$$

$$\frac{1}{\lambda_\infty - \lambda} - \frac{1}{\lambda_\infty - \lambda_0} = \frac{A_0 kt}{\lambda_\infty - \lambda_0} \quad (3)$$

$$A_t = A_0 \left(\frac{\lambda_\infty - \lambda}{\lambda_\infty - \lambda_0} \right) \quad (4)$$

The graphical plot of the relationship between the hydrolysis time (t) towards $(\lambda_\infty - \lambda)$; $\ln(\lambda_\infty - \lambda)$ as well as $1/(\lambda_\infty - \lambda)$ each has a linear regression value (R^2) which varies according to the type of catalyst as described in Figure 8. Based on the linearity values, the usage of homogeneous and activated bentonite catalysts are fitting with the zeroth-order reaction model which meaning the hydrolysis of pineapple leaf fiber was not affected by catalyst concentration. Meanwhile, the presence of inactivated bentonite and the absence of catalyst in the hydrolysis system followed the 1st order reaction model as listed in Table 3.

The hydrolysis of pineapple leaf fibers in this research, with the addition of dilute acid catalyst and the solid acid catalyst, was categorized as a heterogeneous system due to produced higher crystallinity and a lower degree of polymerization of hydro cellulose products. The rate of hydrolysis of cellulose in crystalline form is one to two orders of magnitude lower than that of homogeneous hydrolysis of soluble compounds that can be achieved using a concentrated acid catalyst as listed in Table 4.

Based on the 1st order reaction model, the addition of catalyst particularly activated bentonite was affected in the rate of pineapple leaf fiber hydrolysis and it was also described from zeroth-order reaction even though not quite significant. Meanwhile, the dilute sulfuric acid declined the rate, and on the contrary nitric acid catalysts, raise the rate of the hydrolysis process. The chemical activated bentonite brings a significant rate on hydrolysis reaction and is related to the obtained highest total glucose concentration. The dilute sulfuric acid has a lower rate constant compared to the process with no catalyst due to part of cellulose was re-precipitated and generated a higher hydrolysis rate. A similar result has been reported that 60% dilute acid or less given lower hydrolysis process due to not much-dissolved cellulose so the breaking of β -1-4-glycosidic bonds did not reach the optimum level [20]. To determine the k value of cellulose hydrolysis also required the k value of glucose destruction due to cellulose were not completely degraded to form glucose but also another substance. The k value of glucose destruction was obtained from the slope of chosen order reaction model, where a higher k value meaning higher catalyst concentration and temperature used in the hydrolysis process [25]. From Table 4, according to the 1st order reaction model that the involvement of bentonite catalyst on hydrolysis of pineapple leaf fiber gave higher glucose destruction in line with concentration and temperature used in the system.

4. Conclusion

In summary, pineapple leaf fiber has good prospects for development as a source of lignocellulosic biomass. The involvement of homogeneous and heterogeneous catalysts gives different results both in total sugar content and in the rate of the hydrolysis reaction. In this research, chemical activated bentonite is the most promising catalyst since generating the highest total glucose, meanwhile nitric acid is

Table 4. Reaction rate constant (k).

Type of catalyst system	k_0 (mol.L ⁻¹ .s ⁻¹)	k_1 (s ⁻¹)	k_2 (L.mol ⁻¹ .s ⁻¹)
No catalyst	0.0081	0.0284	0.1434
Inactivated bentonite	0.0050	0.8360	0.1866
Physical activated bentonite	0.0136	0.3240	0.0119
Chemical activated bentonite	0.0097	0.9247	0.1480
Sulfuric acid catalyst	0.0302	0.0080	0.0028
Nitric acid catalyst	0.0023	0.2139	0.0346

not a good choice for assisting the hydrolysis reaction of pineapple leaf fiber. The presence of a heterogeneous catalyst and homogeneous catalyst followed the zeroth-order reaction in the hydrolysis of pineapple leaf fiber.

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